

## GLOSSARY AND CREDITS

So far you have studied measurement, classification and properties of matter, and atomic structure. All of the ideas and concepts you have learned in chemistry are necessary to understand the chemistry you will study in the rest of this course. Scientists will never unlock all of the secrets of nature, but some of the regularities and predictions man has learned will be studied in this unit.

### VOCABULARY

anion                      A negatively charged ion.

cation                     A positively charged ion.

Conservation of Mass A law which states that in ordinary chemical reactions, the sum of the masses of the reactants always equals the sum of the masses of the products.

covalent bond            A chemical bond between atoms with similar electronegativities; a sharing bond; valence electrons are shared.

electronegativity        A measure of the ability of atoms to attract (gain) electrons.

ion                         Any atoms or group of atoms with unequal numbers of protons and electrons, resulting in a net ionic charge.

ion charge                The electron charge an atomic particle carries.

ionic bond                A bond between atoms of greatly differing electronegativities.

metallic bond            The bond formed between metals, characterized by the electrons being shared among all the nuclei.

polar	An unequal, unbalanced distribution of electrons in a molecule causing one portion to be positive and another to be negative.
polarity	The measure of the degree of a charge separation in a polar molecule.
stoichiometry	The study of the quantitative relationships between reactants and products in a chemical reaction.